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Sorption of Pesticide Endosulfan by Electrodialysis Membranes

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Abstract

Endosulfan (ES) is a micropollutant found in reverse osmosis concentrates from water reuse applications. Electrodialysis (ED) can remove and recover charged solutes from such concentrates. While polar compounds cannot normally be removed, their fate in ED is important as they can contribute to membrane fouling/poisoning and be released during cleaning. High adsorption of ES to ED membranes was observed. Consequently, the influence of solution pH and presence of humic acid (HA) on sorption mechanisms of ES to ion-exchange membranes during batch sorption isotherm and ED experiments were investigated systematically. ES-membrane partition coefficients (log K_{EDMEM} (cm^3/mol)) quantified through sorption isotherm experiments suggested that ES sorption was resultant of membrane catalysed ES degradation, hydrogen bonding and cation–π interactions between ES and membrane functional groups. ES sorption at pH 7 (550 µg/cm^3) was greater than sorption at pH 11 (306 µg/cm^3) due to alkaline hydrolysed ES and resultant decrease in bonding capacity with the membranes at high pH. The presence of HA reduced sorption at pH 7 (471 µg/cm^3) and 11 (307 µg/cm^3) due to HA competitive sorption. Partial membrane desorption was noted in isotherm (≤ 20%) desorption experiments and was dependent on the initial mass sorbed, solvent pH and resultant membrane interactions.

Keywords: Electrodialysis; Endosulfan; Adsorption; Ion exchange membranes; Organic matter; Pesticide.

1. Introduction

ES is an organochlorine insecticide that has been used extensively in agriculture for the control of insect pests of crops such as cotton, vegetables, fruit and tobacco [1]. ES movement via runoff and spray drift leads to contamination of aquatic environments [2]. Studies have shown that ES is toxic to fish and aquatic invertebrates [3], while potential human health impacts include cardiovascular, endocrine, gastrointestinal and respiratory toxicity [4]. The European Commission Directive on the Quality of Water Intended for Human Consumption sets a maximum admissible concentration of 0.1 µg/L per individual pesticide and 0.5 µg/L for the sum of pesticides in drinking water [5]. While ES concentrations in aquatic environments are generally low (< 1 µg/L) [2], in areas of application ES has been found in brackish groundwater at 4.2 µg/L [6] and up to 1700 µg/L in surface water [7]. Reverse osmosis (RO) concentrates from water reuse contain concentrated amounts of ES because ES is retained by RO.

ES is a mixture of two stereoisomers, α-endosulfan (α-ES) and β-endosulfan (β-ES) in the range of 64-70% and 29-32%, respectively [8] and may contain 2% endosulfan diol (ES diol) and 1% endosulfan ether (ES ether) [4]. The primary degradation product of concern is endosulfan sulfate (ES sulfate), which is formed in the environment via microbial degradation of α- and β-ES [9]. ES can be hydrolysed to form less toxic ES diol [10].

The treatment of pesticides by membrane processes such as nanofiltration (NF) and RO has been widely reported [11-14]. Two mechanisms for pesticide adsorption to membranes have been suggested: hydrophobic interaction between pesticide and membrane surface [13] and hydrogen bonding between pesticide functional groups and membrane active layer [12]. ED is a viable technology for the desalination of brackish water [15], seawater [16] and solutions containing organics [17]. ED is more cost effective than RO in the salinity range of 1.5-3.5 g/L [18] and has the potential to treat the waste stream generated by pressure driven membrane processes, which are commonly used in desalination and water reuse applications. Such waste streams, or brines, contain salts, nutrients, organic matter (OM) and numerous micropollutants including pesticides, heavy metals, endocrine disrupting chemicals [19].

Mechanisms governing ES behaviour in ED and interaction with ion-exchange membranes have to date not been identified. Previous studies have suggested electrostatic and hydrophobic interactions between organic molecules and ion-exchange membranes [20]. Quantification of these mechanisms has to date not been possible. An understanding of ES partitioning between water and anion- (AEM) and cation- (CEM) exchange membranes is important for prediction of its fate in ED.

Strong interactions between OM and pesticides have been reported with pesticide sorption to OM influencing transport and fate of these compounds in other processes [21-22]. Agbekoko et al. [21] studied atrazine and simazine removal by NF in the presence and absence of OM. Removal increased from 50% to 90-100% with OM. The principle mechanisms for removal in the presence of OM were pesticide-OM complexion, membrane electrostatic repulsion and adsorption. In ED, OM deposits on membranes leading to fouling [23]. However, Park et al. [24] found that a small fraction of OM may permeate AEMs. The influence of OM on ES membrane sorption during ED is unknown.

The objectives of this study were to investigate 1) the interaction mechanisms between ES and (ion-exchange membranes, 2) the influence of solution pH and OM on this interaction and 3) the fate of ES during ED. Membrane-water partition coefficients (log K_{EDMEM}) for ES were determined in sorption isotherm experiments to elucidate the mechanisms governing membrane sorption. The influence of solution pH on ES membrane sorption during batch and continuous ED experiments was evaluated to identify differences in sorption between undissociated (pH 7) and degraded (pH 11) ES. Sorption isotherm and ED experiments were conducted with and without the presence of an OM. Results from this study were compared to membrane sorption of the steroidal hormone estrone during ED, due to similar sorption behaviour, whereby electrostatic interactions along with sorption and diffusion mechanisms were found to play an important role.

2. Materials and Methods

2.1 Chemicals

All chemicals used were analytical grade. NaOH and HCl used for pH adjustments (1 mol/L) and membrane desorption experiments (0.002 mol/L) and Na_2SO_4 (0.5 mol/L) used in the electrode rinse solution, were purchased from Fisher Scientific (UK). All experiments were conducted in a background solution of 5 g/L NaCl and 84 mg/L NaHCO_3 (Fisher Scientific, UK). Radiolabeled [2,3-3^14]C ES (> 95% purity; 18.5 MBq solid form) was purchased from the Institute of Isotopes Co., Ltd. (Hungary). Radiolabeled [2,4,5,7,3^3^14]H estrone (> 98.5% purity; 37 MBq/mmol) was purchased from Perkin Elmer (UK). 37 MBq/mL (5.5 TBq/mmol)
was purchased from GE Healthcare (UK). Non-labelled ES compounds (≥ 99% purity) and estrone (≥ 98% purity) were purchased from Sigma Aldrich (UK). Stock solutions of radiolabeled (100 µg/L) and non-labelled (990 ng/mL and 1000 ng/mL) ES compounds and estrone were prepared in methanol (CH\textsubscript{3}OH) (Fisher Scientific, UK). Dichloromethane (CH\textsubscript{2}Cl\textsubscript{2}), used in the liquid-liquid extraction of unlabelled ES compounds was obtained from Acros Organics (Belgium) while trifluralin (75 mg/L) was used as an internal standard (Riedel-de Haen, Germany). Physicochemical properties of the ES isomers, degradation products and the hormone estrone are outlined in Table 1.

The OM selected for use was humic acid sodium salt (HA) (Sigma Aldrich, UK). Although OM concentration in natural waters is variable (0.5-100 mg C/L) [23], a concentration of 12.5 mg C/L was used for experiments containing HA. The major functional groups of the negatively charged (neutral basic pH) HA include carboxylic, phenolic and alcohol/phenol/amides, and methoxy [26].

### 2.2 Sorption isotherm and desorption experiments

Radiolabeled (100 ng/L) and non-labelled ES (70-30 α-ES/β-ES mix) was added to 100 mL background solution (adjusted to pH 7) to make up concentrations of 1, 2, 10, 100 and 2500 µg/L. A 2 cm\textsuperscript{2} segment of AEM or CEM (3 replicates/membrane) was added to each solution bottle and shaken in an incubator shaker (Certomat BS-1, Sartorius) at 200 RPM and 25°C for 100 hours. A sample (1 mL) was taken from each bottle prior to membrane addition and periodically during each experiment for ES analysis (total 11 samples per bottle). Solutions without membrane addition were included to determine possible ES sorption to the glass bottles. For experiments undertaken in the presence of HA the bottles were shaken for 24 hours prior to membrane addition to allow for ES-HA equilibrium. Samples (3.5 mL) from each bottle were then taken for analysis and periodic during each ES-HA experiment for HA analysis. ES desorption from membranes used in the 1 µg/L isotherm experiment was measured in 100 mL solutions of 0.002 mol/L NaOH/HCl/ultrapure water shaken for 170 hours.

### 2.3 Electrodialysis system, membranes and protocol

A BEL-500 stack (Berghof, Germany) with six Neosepta AMX-SB AEMs and seven CMX-SB CEMs (supplied by Eurodia, Germany; manufactured by ASTOM Corporation, Japan) with an available membrane area of 58 cm\textsuperscript{2} each was used in the ED experiments. The positively charged AEMs (basic) and negatively charged CEMs (acidic) used in the sorption isotherm and ED experiments contain alkylammonium and sulfonic acid ion-exchange groups, respectively, attached to the aromatic rings of a polystyrene-divinybenzene (PS-DVB) matrix on a microporous polystyrene chloride (PVDC) gel supported by plasticized PVC cloth [27]. The thicknesses of the membranes were 0.14 and 0.17 mm for the AEM and CEM, respectively. The volumes of the membranes used during ED experiments were 4.9 cm\textsuperscript{3} and 6.9 cm\textsuperscript{3} for the 58 cm\textsuperscript{2} and 6.9 cm\textsuperscript{2} membrane systems, respectively [28]. The stack was connected to a DC electric potential (GW Instek DC Power supply Model GPR-1810HD, Taiwan) through TiO\textsubscript{2}-coated titanium electrodes. Flow rates in all compartments (diluate, concentrate and electrode rinse) were 1.5 L/min (Masterflex I/P Variable speed pump system, USA). Continuous (diluate and concentrate recirculated to one feed container) and batch (separate diluate and concentrate) were undertaken.

An applied voltage of 10 V was fixed for all ED experiments. Preliminary batch ED experiments (4L each, separate diluate and concentrate) with 100 µg/L (non-labelled) α-ES and ES sulfate were undertaken to establish the behaviour of these compounds during ED at pH 3 and 7. Feed solutions (2500 µg/L ES, 4L total) for subsequent continuous ED experiments were prepared in bagged solution. The ES concentration used in the ED experiments was much greater than the concentration usually found in natural waters due to the high sorption capacity of the membranes. To determine the influence of solution pH on ES sorption during continuous ED experiments, the feed pH was maintained constant by the addition of 1 mol/L HCl and/or NaOH. A new set of membranes was used in each continuous ED experiment to avoid interference from ES sorbed in previous experiments. Sorption within the diluate and concentrate was evaluated in batch experiments undertaken after the completed continuous experiments (continuous solution separated into diluate and concentrate). A mixed ES and estrone (2500 µg/L each) continuous experiment was undertaken to evaluate competitive sorption between compounds exhibiting similar behaviour. Samples were collected at the beginning of each ED experiment and periodically for ES and/or estrone (1 mL) and UV-Vis absorbance (3 mL) analysis.

### 2.4 Analytical methods

Radiolabeled ES and estrone were analysed using a Beckman Coulter scintillation counter (LS 6500). Samples (1 mL) were mixed with 7 mL Ultima Gold\textsuperscript{®} scintillation cocktail (Perkin Elmer, UK) and measured in triplicate. ES and estrone concentration was ascertained from a linear regression performed on calibration standards (ES: 0.1, 1, 2 and 5 µg/L; Estrone: 0.01, 0.1, 1, 10, 100 and 1000 ng/mL). Gas Chromatography-Electron Capture Detection (GC-ECD) (Autosystem XL chromatograph with an electron-capture detector, Perkin Elmer, UK) was used for the determination of unlabelled α-ES, ES sulfate and ES diol within samples (100 mL) taken during preliminary ED experiments. Liquid-liquid extraction was undertaken with dichloromethane (50 mL) and the internal standard trifluralin (5 µL).

The pH, electrical conductivity and temperature of samples periodically taken from the feed, diluate and concentrate during ED experiments were measured using a pH/Conductivity meter (Multiline P4 pH electrode, WTW, Germany). UV-Visible Spectrometer (Varian Cary 100 Scan, UK) was used to determine the absorbance of HA (wavelength of 254 nm) in experimental samples.

### 3. Results and discussion

#### 3.1 Preliminary electrodialysis experiments

In the early stages of this research, significant difficulties were encountered with the ES mass balance. Preliminary ED experiments showed a decrease in α-ES and ES sulfate within the diluate and concentrate (Figure 1). The concentration of α-ES within the diluate and concentrate (91.4% and 88.6% initial mass lost, respectively) decreased more than that of ES sulfate (80.6% (diluate) and 79.1% (concentrate)) at pH 3. The same trend was observed at pH 7. This observation led to thorough investigations of (a) losses to equipment, (b) volatilisation, and (c) degradation in an attempt to recover the ES.

![Figure 1](image-url)

Losses to equipment (a) are material dependent. Polystyrene sample containers were reported to sorb ES strongly, compared with glass [29] and stainless steel. In consequence, all process stream lines were converted to plastic to stainless steel where possible (small sections of silicon tubing were required for the peristaltic pump). ED experiments with these materials and stainless steel where possible (small sections of silicon tubing were required for the peristaltic pump). ED experiments with stainless steel were undertaken to establish the behaviour of ES under similar conditions.

Volatilisation (b) of ES from aqueous media in laboratory experiments has been attributed to its low water solubility and high volatility [29]. Control sorption experiments were undertaken to measure ES sorption to and/or volatilisation from 100 mL glass sample bottles. A loss of up to 20% was noted with the bulk lost in the beginning of the experiment. Log K\textsubscript{OM/EDM} values determined in the sorption isotherm experiments were adjusted accordingly to account for this loss due to volatilisation. Diluate and concentrate containers were subsequently closed to prevent volatilisation during the experiments.

Heterogeneous electrochemical degradation (c) of organic compounds may occur on membranes and electrodes [30]. To establish if this is the case, diluate and concentrate were analysed for the presence of the degradation product ES diol (Figure 1A and C). ES diol was not detected in the initial diluate and concentrate samples at pH 3. However, a slight increase to 10.1 and 9.5 µg
Detection of ES diol (but not individual identification) in the form of ES was possible because of surface. It is postulated that ES is hydrolysed to ES diol on contact with the strongly basic AEM. This could further explain the higher sorption of ES to the AEM membranes. Bonding between ES and the CEM would be facilitated through interaction of the non-metallic cation RNH$^+$ with double bonds can be thought of as a form of X-H$^\alpha$ hydrogen bonds. It is difficult to define the contribution to bonding from a specific interaction due to the fact that many different forces (e.g. donor-acceptor, cation-$\pi$, hydrophobic interactions, dispersion and van der Waals forces) contribute to molecular binding. However, the occurrence of cation-$\pi$ interactions could further explain the higher sorption of ES to the ion-exchange membranes. Bonding between ES and the CEM would be facilitated between the ES S=O group (C-7) and the CEM CH$_3$ hydrogen-donor group, while bonding between ES diol and the AEM would be between the C-7 and C-10 OH groups. However, the AEM quaternary ammonium group is dissociated and the cation may be strongly hydrated. Therefore, a novel interaction mechanism not previously considered is proposed which is cation-$\pi$ interactions [38]. The fixed positive charges of the AEM have high potential to attract the negative electron imbalance of the ring structure. The interaction of the non-metallic cation RNH$.^+$ with double bonds can be thought of as a form of X-H$^\alpha$ hydrogen bonds. It is difficult to define the contribution to bonding from a specific interaction due to the fact that many different forces (e.g. donor-acceptor, cation-$\pi$, hydrophobic interactions, dispersion and van der Waals forces) contribute to molecular binding. However, the occurrence of cation-$\pi$ interactions could further explain the higher sorption of ES to the ion-exchange membranes. Bonding between ES and the CEM would be facilitated between the ES S=O group and the CEM OH group. In aqueous solutions, ES diol is the most common degradation product and can be reduced on a charged surface. It is postulated that ES is hydrolysed to ES diol on contact with the strongly basic AEM. Detection of ES diol (but not individual identification) in the form of ES was possible because of the utilisation and position of the $^{13}$C label on ES. Bonding between ES diol and the CEM functional group would be facilitated through the ES diol OH bipolar groups.

At pH 7, when HA is $> 99\%$ dissociated, hydrogen bonding between the ES S=O and C-O groups and HA OH groups would occur. In the presence of HA ES sorption at pH 7 decreased by 79 µg/cm$^3$ (471 µg/cm$^3$), Figure 4) due to electrostatic repulsion between negatively charged ES-HA complexes and the CEMs. ES sorption at pH 11 was similar in the presence of HA (307 µg/cm$^3$, 33.6%) with the kinetics indicating membrane diffusion. After the feed solution was separated, sorption continued within the diluate and concentrate at pH 7; indicating sorption to both the AEMs and CEMs. This confirms the results by Pronk et al. [33] whereby neutral compounds were assumed to sorb to both AEMs and CEMs. Above pH 10, ES degradation results from cleavage of the cyclic sulfite group [10] leading to the formation of ES diol. This process is accelerated by NaOH addition. The reduced sorption of ES diol at pH 11 can be predominantly attributed to the reduced bonding capability between ES diol and the membranes compared to ES.

At pH 11, the only other possible mechanism for ES losses is sorption to the membrane polymers. Sorption isotherm experiments were undertaken to quantify such ES sorption to the ion-exchange membranes.

3.3 Endosulfan sorption mechanisms

Prón et al. [33] studied the sorption of the hormone 17a-ethinylestradiol (75% initial mass sorbed) to ion-exchange membranes during batch ED experiments for the treatment of urine and postulated that sorption was due to hydrophobicity. However, a study by Banasik [34] suggested that trace organic sorption (of steroidal hormones) to ion-exchange membranes was influenced by mechanisms other than hydrophobic interactions. The possible interactions between ES and the AEM and CEM functional groups are illustrated in Figure 3. Previous studies have demonstrated that trace organics primarily interact with membranes through hydrogen bonding [35-36]. The AEM functional group is capable of forming hydrogen bonds with molecules containing hydrogen-donor and acceptor groups [37] and presents more opportunities for bonding; thus accounting for greater ES sorption affinity compared to the CEM. AEM sorption of ES would be facilitated through bonding between the S-1 (S=O, C-7) and C-10 (C-O) hydrogen-acceptor groups and the AEM CH$_3$ hydrogen-donor group, while bonding between ES diol and the AEM would be between the C-7 and C-10 OH groups. However, the AEM quaternary ammonium group is dissociated and the cation may be strongly hydrated. Therefore, a novel interaction mechanism not previously considered is proposed which is cation-$\pi$ interactions [38]. The fixed positive charges of the AEM have high potential to attract the negative electron imbalance of the ring structure. The interaction of the non-metallic cation RNH$.^+$ with double bonds can be thought of as a form of X-H$^\alpha$ hydrogen bonds. It is difficult to define the contribution to bonding from a specific interaction due to the fact that many different forces (e.g. donor-acceptor, cation-$\pi$, hydrophobic interactions, dispersion and van der Waals forces) contribute to molecular binding. However, the occurrence of cation-$\pi$ interactions could further explain the higher sorption of ES to the ion-exchange membranes. Bonding between ES and the CEM would be facilitated between the ES S=O group and the CEM OH group. In aqueous solutions, ES diol is the most common degradation product and can be reduced on a charged surface. It is postulated that ES is hydrolysed to ES diol on contact with the strongly basic AEM. Detection of ES diol (but not individual identification) in the form of ES was possible because of
trace organics such as ES can desorb into the diluate during reversal or cleaning. The significant desorption observed in this study further implies that wastewater obtained from the cleaning process of ED membranes used to treat water containing pesticides could contain potentially high concentrations of such contaminants. While the US EPA recommends less than 74 µg/L ES in waterways, exposure to just 1.3 µg/L is toxic to fish [3]. Exposure for short periods can cause adverse nervous system, stomach, blood, liver and kidney effects and death. Further treatment of the diluate, concentrate and cleaning solutions need to be considered when trace organics are amongst the target contaminants. These results are important as the risks associated with the disposal of ED membrane cleaning solutions have been not yet been addressed.

3.6 Comparison with hormone sorption

To gain insight into the general applicability of observations to other trace organics, the sorption of ES during ED was compared to that of estrone (E1). Estrone is a natural estrogen that is dissociated and negatively charged above pH 10.4 [42]. Estrone was selected due to its presence in effluents from conventional wastewater treatment plants [43] and its high sorption affinity for the ion-exchange membranes [34]. A dual compound ED experiment with both ES and estrone was undertaken at pH 7 and 11. Preferential sorption of ES was observed (Figure 5). Sorption ES and estrone was reduced at pH 11 possibly because of competitive sorption between the degraded/dissociated ES and estrone and the ion-exchange membranes. It is interesting to note that ES sorption was greater in the single compound experiment (pH 7: 590 ± 17.9 µg/cm²; pH 11: 414 ± 12.4 µg/cm³) than in the single compound experiment (pH 7: 550 ± 16.5 µg/cm²; pH 11: 306 ± 9.2 µg/cm²). A similar trend was noted for estrone at pH 11 (Single: 257 ± 4.9 µg/cm²; Dual: 278 ± 5.6 µg/cm²). However, the mass of estrone sorbed in the single compound experiment (pH 7: 380 µg/cm²) was similar to the mass sorbed in the dual compound experiment (pH 7: 381 µg/cm³). While further work is required to fully understand this competitive behaviour, results indicate that the sorption capacity of ED membranes is very high. In consequencesaturation with ES and estrone was not reached at conditions that reflect the upper end of environmental concentrations.

[Figure 5]

4. Conclusions

The quantification of partition coefficients indicated strong sorption of ES to the ion-exchange membranes. While challenges remain to elucidate the contribution of specific interaction mechanisms to this sorption, it was postulated for those to occur predominantly due to hydrogen and cation-π interactions between ES and membrane functional groups. Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups, Alkaline hydrolysis of ES to ES diol influenced sorption whereby lower hydrogen bonding capacity between ES diol and the membranes reduced sorption at pH 11. The influence of solute-solute interactions between ES and membrane functional groups.

Acknowledgements

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Appendices

For further information on the calculation ES sorption to each membrane and the water-membrane partition coefficients along with sorption/desorption kinetics during batch isotherm and ED experiments and details of steroidal hormone sorption, refer to Appendix A.

References

Table Captions

Table 1. Physicochemical characteristics of endosulfan compounds and the hormone estrone.

<table>
<thead>
<tr>
<th>Property</th>
<th>Endosulfan</th>
<th>Estrone</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>α</td>
<td>β</td>
</tr>
<tr>
<td>Formula</td>
<td>C₉H₆Cl₆O₃S</td>
<td>C₉H₆Cl₆O₃S</td>
</tr>
<tr>
<td>Molecular Weight (g/mol)</td>
<td>406.93</td>
<td>406.93</td>
</tr>
<tr>
<td>Solubility in water (mg/L 20°C) [44-45]</td>
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<td>0.45</td>
</tr>
<tr>
<td>Log Kₘₐₐₙ [47]⁺</td>
<td>3.83 (3.50)</td>
<td>3.83 (3.50)</td>
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<tr>
<td>Hydrogen acceptors</td>
<td>3</td>
<td>3</td>
</tr>
<tr>
<td>Hydrogen donors</td>
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<td>Structure b</td>
<td><img src="image1.png" alt="Structure" /></td>
<td><img src="image2.png" alt="Structure" /></td>
</tr>
</tbody>
</table>

b Values inside parentheses are estimated using KowWin Log P software [48]; 

b Asterix on α–ES indicates location of ¹⁴C radiolabel.
Figure Captions

Figure 1. Concentration (µg/L) of α-ES, ES sulfate and ES diol within the concentrate (A, B) and dilute (C, D) at pH 3 and pH 7 (Average initial concentration of α-ES and ES sulfate were 92.4 and 92.0 µg/L, respectively).

Figure 2. ES sorption isotherms for the ion exchange membranes (A) CEMs and (B) AEMs with and without HA (± S.D. given to indicate standard deviation associated with data; time used to determine log $K_{AEM/CEM}$: 100 h).

Figure 3. Schematic of possible interactions between ES molecules and the CEM and AEM functional groups (Functional groups attached to polystyrene-divinylbenzene (PS-DVB) copolymer matrix).

Figure 4. Sorption (µg/cm$^3$) of ES to ion-exchange membranes during ED experiments (Average initial mass of ES 11176 ± 317 µg).

Figure 5. Sorption (µg/cm$^3$) of ES and estrone during mixed ED experiments (Initial mass of ES and estrone in solution were 889 ± 13 µg/cm$^3$ and 904 ± 37 µg/cm$^3$, respectively).
Figure 2

- A - CEM
- B - AEM

\[
\begin{align*}
\log C_{\text{AEM/CEM}} (\text{ng/cm}^3) &\quad \log C_w (\text{ng/L}) \\
\end{align*}
\]

Without HA
HA

Figure 3

- Anion exchange membrane (AEM)
- Cation exchange membrane (CEM)
- Alkylammonium functional group
- Sulfonic acid functional group
- ES-diol
- ES (α-ES + β-ES)

Hydrogen acceptor
Hydrogen donor
Bipolar
Figure 4

Figure 5

Appendix A

Sorption of Pesticide Endosulfan by Electrodialysis Membranes
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Overview of Appendix A
In this Appendix we present:
1. Endosulfan sorption kinetics during 1 µg/L endosulfan sorption isotherm experiments;
2. Mathematical equations related to the determination of the mass of endosulfan sorbed to membranes during batch isotherm experiments;
3. Information regarding the determination of the water-membrane partition coefficients (log $K_{AEM/CEM}$);
4. Endosulfan desorption kinetics during 0.002 M HCl, NaOH and ultrapure water desorption experiments;
5. Details on trace organic sorption of steroidal hormones to ion-exchange membranes from study by Banasiak [34];
6. Figure showing the concentration of endosulfan within the dilute and concentrate at pH 7 and 11 during batch ED experiments.

All information found in this Appendix also referred to in the manuscript.

Endosulfan sorption kinetics during 1 µg/L endosulfan sorption isotherm experiment

![Figure A1](image)

Figure A1. ES sorption kinetics for the AEM and CEM (1mM NaHCO3, 85.5 mM NaCl, pH 7, 1 µg/L ES; Vertical dotted line indicates time chosen for log $K_{AEM/CEM}$ determination).

Determination of mass of endosulfan sorbed to membrane during batch isotherm experiments
The mass of ES sorbed to each membrane ($m_{AEM/CEM}$, µg or ng) was calculated using mass balance eqn (A1).

$$m_{AEM/CEM} = m_e^w - m_e^t = C_e^w V_e^w - C_e^t V_e^t$$

where, $m_e^w$ and $m_e^t$ are the initial (superscript 0) and $t = 100$ hr (superscript e) mass of ES freely dissolved in aqueous solution (subscript w) (µg or ng), $C_e^w$ and $C_e^t$ are the initial and $t = 100$ hr ES concentration (µg/L or ng/L), $V_e^w$ and $V_e^t$ are the initial and T = 100 hr solution volumes (L).

Equilibrium concentrations of ES sorbed per unit volume of membrane ($C_{AEM}$ or $C_{CEM}$, µg/cm$^3$ or ng/cm$^3$) were calculated using eqn (A2).

$$C_{AEM/CEM} = \frac{m_{AEM/CEM}}{V_{AEM/CEM}} = \frac{(C_e^0 - C_e^t) V_e}{V_{AEM/CEM}}$$

(A2)

where, $V_{AEM/CEM}$ is the AEM or CEM membrane volume (cm$^3$).

Partition Coefficient ($K_{AEM/CEM}$) Determinations
The partition coefficient, $K_{AEM/CEM}$ (L/cm$^3$) for ES between the respective membrane (AEM or CEM) and the bulk solution was evaluated using eqn (A3).

$$K_{AEM/CEM} = \frac{C_{AEM/CEM}}{C_{CEM/CEM}} = \frac{m_{AEM/CEM}}{m_{CEM/CEM}} V_{CEM/CEM}$$

(A3)

where $C_{AEM/CEM}$ is the concentration of ES sorbed per unit volume of membrane at time $t = 100$ hr (µg/cm$^3$ or ng/cm$^3$), $C_e^w$ is the ES concentration (µg/L or ng/L), $m_{AEM/CEM}$ is the mass of ES sorbed to each membrane (µg or ng), $V_{AEM/CEM}$ is the AEM or CEM membrane volume (cm$^3$), $V_e$ is the solution volume (L) and $m_e^w$ is the mass of ES freely dissolved in aqueous solution (subscript w) (µg or ng).

Log $K_{AEM/CEM}$ can be derived from the slope of the linear regression of $C_{AEM/CEM}$ as a function of $C_e^w$, if the sorption isotherms (plotted on a log scale) are linear. Since the concentration range is several orders of magnitude, the sorption isotherms were plotted on a log scale (eqn (A4)) and the slope of the regression $n_i$ indicates the sorption isotherm deviation from linearity [32].

$$\log C_{AEM/CEM} = \log K_{AEM/CEM} + n_i \log C_e^w$$

(A4)

As $n_i$ was close to 1 for all isotherms, this was set to 1, which is equivalent to the linear relationship described by eqn (A3). Standard deviation (S.D.) associated with log $C_{AEM/CEM}$ values were calculated and presented as ± S.D. Confidence intervals of the mean associated with the log $C_{AEM/CEM}$ and Log $K_{AEM/CEM}$ values were also calculated and presented as ± C.I.
Endosulfan desorption kinetics during 0.002 M HCl, NaOH and ultrapure water desorption experiments

![Endosulfan desorption kinetics figure]

Figure A1. ES desorption kinetics for the AEM and CEM (1mM NaHCO₃, 85.5 mM NaCl, 0.002 M HCl, 0.002 M NaOH and ultrapure water (UW)).

Details on trace organic sorption of steroidal hormones to ion-exchange membranes from study by Banasiak [34]

The mechanisms of sorption of four steroidal hormones – estradiol, estrone, progesterone and testosterone – to ED membranes were investigated as a function of solution pH and presence of HA. Water-membrane partition coefficients \( \log K_{AEM/CEM} \) were determined through batch sorption isotherm experiments using the same methods as this present study. Although Pronk et al. [33] postulated that hormone sorption to ion-exchange membranes was related to hydrophobicity, poor correlation between hormone \( \log K_{ow} \) and \( \log K_{AEM/CEM} \) values suggested that other mechanisms, such as electrostatic interactions and hydrogen bonding, were important in hormone sorption at neutral pH.

Concentration (µg/L) of endosulfan in the diluate and concentrate at pH 7 and pH 11 in batch ED experiments

![Concentration of endosulfan figure]

Figure A2. Concentration (µg/L) of ES in the diluate and concentrate at pH 7 and pH 11 in batch ED experiments without humic acid (1 mM NaHCO₃, 85.5 mM NaCl, 10 V; ES initial concentration: pH 7 1188 µg/L, pH 11 1670 µg/L).