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Citation for published version:

Digital Object Identifier (DOI):
10.1063/1.1378042

Link:
Link to publication record in Edinburgh Research Explorer

Document Version:
Publisher's PDF, also known as Version of record

Published in:
Journal of Chemical Physics

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Comparison of near-threshold reactivity of ground-state and spin-orbit excited chlorine atoms with methane

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Citation: J. Chem. Phys. 115, 179 (2001); doi: 10.1063/1.1378042
View online: http://dx.doi.org/10.1063/1.1378042
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Comparison of near-threshold reactivity of ground-state and spin-orbit excited chlorine atoms with methane

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(Received 28 February 2001; accepted 18 April 2001)

A 4:1 mixture of CH4 and BrCl diluted in He are coexpanded into a vacuum chamber and the reaction of methane with atomic chlorine is initiated by photolysis of BrCl. Near 420 nm, the resulting mixture of ground- and excited-state chlorine atoms have spatial anisotropies of $\beta_{\text{phot}} = -0.7$ for the Cl$(^2P_{3/2})+\text{Br}$ channel and $\beta_{\text{phot}} = +1.8$ for the Cl$(^2P_{1/2})+\text{Br}$ channel. The speed-dependent spatial anisotropy $\beta_{\text{tran}}(\nu)$ of the CH3$(\nu=0)$ reaction product is detected by 2 + 1 resonance-enhanced enhanced multiphoton ionization. Our results indicate that the Cl* + CH4 reaction is unimportant in the near-threshold collision energy range of 0.13–0.16 eV, whereas the reaction with ground-state Cl atoms with CH4 excited with one quantum in the $v_2$ (torsion) or $v_3$ (bending) mode is dominant. © 2001 American Institute of Physics. [DOI: 10.1063/1.1378042]

I. INTRODUCTION

The reaction Cl + CH4 → HCl + CH3 has drawn much attention from the environmental, experimental, and theoretical points of view. In addition to being an important reaction in combustion at high temperatures, this reaction is believed to be one of the termination steps of the ozone destruction cycle in the stratosphere. The modeling of these processes requires accurate parametrization of rate constants at wide temperature ranges, and has prompted detailed kinetics studies.1–5 This reaction shows pronounced non-Arrhenius behavior at both low and high temperatures. Among other possible sources for this behavior, Ravishankara and Wine3 postulated that the reactivity of spin-orbit excited Cl$(^2P_{3/2})$ is enhanced over that of ground-state Cl$(^2P_{3/2})$ and that this enhancement is mostly responsible for the non-Arrhenius behavior below 240 K. This postulate is based on the assumption that the extra spin-orbit energy (881 cm$^{-1}$) of Cl* is available for overcoming the reaction barrier, causing Cl* to be more reactive than Cl.

Often, Cl* reactivity has been believed to be the cause of disparities between the early kinetics data, but the proposed reaction, Cl$(^2P_{3/2}) + \text{CH4} \rightarrow \text{HCl}(X^1\Sigma^-) + \text{CH}_3(X^2A')$, is symmetry forbidden. It becomes possible only via a nonadiabatic transition between two different potential energy surfaces. Recent studies1,4 show that the cross section for the Cl* quenching by CH4 (Cl* + CH4 → Cl + CH3) is larger by more than 2 orders of magnitude than the reaction cross section for the ground-state reaction Cl + CH4 → HCl + CH3. Therefore, for conventional kinetics experiments conducted under multiple-collision conditions, it can be argued that the role of Cl* reactivity is difficult to assess unless the Cl* reaction cross section is large enough to compete with the facile Cl* quenching channel. Matsumi et al.1 attempted to measure the Cl* + CH4 reaction rate indirectly, and they have given an upper bound of 30% of total Cl* removal (Cl* reaction + quenching) rate. Nevertheless, the relative reactivity of Cl* + CH4 compared with Cl + CH4 reactivity has not yet been measured.

A few experiments concerning the reactivity of spin-orbit excited halogen atoms ($X^*$) have been reported for the related reactions $X^* + \text{H}_2 \rightarrow \text{HX} + \text{H}$, where $X=F, \text{Cl},$ and Br. Lee and Liu6 observed that different methods for generating Cl atoms (photolysis of Cl2 vs discharge of Cl2) led to slightly different differential cross sections in their study of the Cl + H2 → HCl + H reaction. They attributed the difference to the enhanced reactivity of Cl* over Cl. Nizkorodov et al.8 studied the HF product state distribution of F + H2 → HF + H at various collision energies. They noticed that part of the product distribution could not be explained by F$(^2P_{3/2})$ reacting with H2$(\nu = 0)$, but could be rationalized by assuming the reaction with F$(^2P_{1/2})$ occurs. Rigorous theoretical calculation of the effect of nonadiabatic interaction in a bimolecular reaction remains a challenging problem, even for simple chemical systems.9

In this paper, we report the first comparison of Cl and Cl* reactivity with CH4 near the reaction threshold. We use what is called the photoloc (photo-initiated bimolecular reaction by law-of-cosines) technique.10 The spatial anisotropy of the CH3 product is measured and compared with the expected anisotropy for Cl* and Cl reacting with methane, thereby allowing us to directly assess the relative reactivity.

Figure 1 shows the energetics involved in Cl/Cl* + CH4 → HCl + CH3 reaction together with the total (translational + internal) energies involved in this work. The Cl atom reaction is endothermic by 600 cm$^{-1}$,11 whereas Cl* reaction is exothermic by 281 cm$^{-1}$, owing to the 881 cm$^{-1}$ difference of the spin-orbit energy of Cl* compared to Cl. The activation barrier for the Cl–atom reaction, as derived from the Arrhenius fit of reaction rate at temperatures between 200 and 300 K, is 862 cm$^{-1}$,11 and the most recent

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semiempirical potential energy surface by Corchado et al.\textsuperscript{12} predicts a barrier height of 1574 cm\(^{-1}\), including the zero-point energies of reagents and products. It should be noted that the experimental activation barrier from the low-temperature rate measurement underestimates the actual barrier height because of the non-Arrhenius behavior at low temperature. As such, the theoretical reaction barrier is usually larger than the activation energy. The calculations\textsuperscript{12–15} indicate: (1) the transition state is collinear along the Cl–H–C axis; (2) the reaction has a ‘late’ barrier located in the exit valley; and (3) the reaction rate is enhanced by C–H stretching mode excitation of CH\(_4\). The series of experiments by Zare and co-workers\textsuperscript{10,16–18} agree with these results.

We employed photodissociation of BrCl near 420 nm as a source of Cl and Cl* atoms. Photodissociation of BrCl at this wavelength involves two major pathways: the Cl+Br channel via C\(^1\Pi(1)\)–X\(^1\Sigma(0^+)\), and the Cl*+Br channel mostly via B\(^3\Pi(0^+)\)–X\(^1\Sigma(0^+)\) transitions. The relative cross section of Cl*+Br vs Cl+Br and the spatial anisotropies at various photolysis wavelengths have been characterized by Cao et al.\textsuperscript{19} and by Cooper et al.\textsuperscript{20} The yield of Cl* changes from 37% to 57% as the photodissociation wavelength varies from 410 to 430 nm. The spatial anisotropy of each channel has a different value and changes slightly across the wavelengths used. Although the center-of-mass translational energy (E\(_{\text{coll}}\)) is slightly smaller for the Cl* reaction based on energy conservation, the total energy available for the reaction is larger for the Cl* reaction for a given photolysis wavelength of the BrCl molecule (see Fig. 1).

The spatial anisotropy of the reaction product \(\beta_{\text{rxn}}(\nu)\) from a bimolecular reaction initiated by a photolysis process that is characterized by a spatial anisotropy \(\beta_{\text{phot}}\) is determined by

\[
\beta_{\text{rxn}}(\nu) = \beta_{\text{phot}}P_2(\cos \alpha),
\]

where \(P_2\) is the second-order Legendre polynomial and \(\alpha\) is the angle between the center-of-mass velocity and the product velocity vectors, as determined by the kinematics of the reaction. From Eq. (1), it is clear that the laboratory velocity resolved (hence angle \(\alpha\) resolved) spatial anisotropy of the particular product is directly proportional to the spatial anisotropy of the photolysis step. The prediction of spatial anisotropy however does not require knowledge of the differential cross section of the product. In this experiment, we have two dissociation channels (Cl*+Br and Cl*+Br) for the photolysis step. The measured spatial anisotropy of the CH\(_3\) product is an average of two spatial anisotropies of the CH\(_3\) product that are associated with each channel in the photolysis step, being weighed by the relative reactivity of Cl vs Cl* and by the relative flux of Cl vs Cl*. Therefore, with a knowledge of the relative yield for Cl and Cl* and the spatial anisotropies for Cl and Cl* from the photolysis, it is possible to obtain the relative reactivity of Cl and Cl* with methane by measuring the CH\(_3\) product spatial anisotropy.

II. EXPERIMENT

Details of the experimental setup and techniques have been discussed previously.\textsuperscript{10,21} BrCl was synthesized in a glass bulb by mixing Cl\(_2\) and Br\(_2\) with 1:3 ratio, and waiting for >30 min. Br\(_2\) (Aldrich, 99.5\% stated purity) is degassed by several freeze–pump–thaw cycles with liquid nitrogen and dry ice before mixing with Cl\(_2\) (Matheson, 99.999%). Because the equilibrium constant for Br\(_2\)+Cl\(_2\)↔2BrCl is only 7 at room temperature,\textsuperscript{20} separation of BrCl from unreacted Br\(_2\) and Cl\(_2\) is not feasible. Unreacted Cl\(_2\) and Br\(_2\) can be photolyzed by 420 nm to produce Cl, Cl*, Br, and Br* atoms. However, the photodissociation cross section of Cl\(_2\) at this wavelength is minimal compared with the cross section of BrCl, and the kinetic energy of Cl atoms produced from Cl\(_2\) at this wavelength is too low to overcome the endothermicity (660 cm\(^{-1}\)) to generate product. Also, we confirmed experimentally that neither the presence of Br\(_2\) nor Cl\(_2\) produced reaction signal, by running control experiments with mixes of Br\(_2\)/CH\(_4\)/He and Cl\(_2\)/CH\(_4\)/He. The BrCl was further mixed with CH\(_4\) (Matheson, 99\%) and He (Liquid Carbonic, 99.995\%) with a ratio of BrCl:CH\(_4\):He=1:4:5 and delivered to the pulsed nozzle (General valve series-9, 0.6 mm orifice size) where it was supersonically expanded into the ionization region of a core-extracted Wiley–McLaren type time-of-flight (TOF) spectrometer. The reaction was initiated by a photolysis beam (410–430 nm) that was produced by frequency doubling (by a BBO crystal) the Nd:YAG laser pumped dye laser output (PL9020 and ND6000, Continuum; LDS867, Exciton). The CH\(_3\) products were allowed to accumulate for

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**FIG. 1.** Energetics of the reaction of Cl/Cl* with CH\(_4\). Cl\(^2\Pi(3/2)\) + CH\(_4\) correlates adiabatically to HCl+CH\(_3\)\(^2\Sigma\) (solid line). Cl\(^2\Pi(3/2)\) + CH\(_4\) correlates to HCl+CH\(_3\)\(^2\Pi\) that is energetically inaccessible with the collision energies used (solid line). The reaction Cl\(^2\Pi(1/2)\) + CH\(_4\) →HCl+CH\(_3\)\(^2\Delta\) requires nonadiabatic transition between the potential energy surfaces (shown in dashed line, effective reaction barrier height is arbitrarily drawn). The total energies are measured from the bottom of Cl+CH\(_4\) (\(\nu=0\)), and their spread for Cl+CH\(_4\) (solid curve) and Cl* + CH\(_4\) (shaded curve) at \(E_{\text{coll}}=0.14\) eV are represented by the Gaussian distributions that would result from a estimated beam translational temperature of 15 K (see Ref. 29).

Note that although the center-of-mass translational energy of Cl+CH\(_4\) is slightly larger than that of Cl\(^*\) + CH\(_4\) for a given wavelength of photolysis, the total energy available for the reaction is higher for the Cl\(^*\) + CH\(_4\).
60–100 ns before being ionized by 2 + 1 resonantly enhanced multiphoton ionization (REMPI) through the \( \tilde{3}p_2 \tilde{2}A''\rightarrow \tilde{X} \tilde{2}A'' \) transition. Frequency-doubled light near 333.3 nm from a Nd:YAG pumped dye laser system (DCR-2A, Spectra-Physics; FL2002, Lambda Physik; DCM and LDS698, Exciton) drove the REMPI process. The resulting ions were detected with microchannel plates. The reaction products were distinguished from background signals through time-jump subtraction. A photoelastic modulator (PEM-80, Hinds International Inc.) flipped the direction of the linear polarization of the photolysis laser beam between parallel and perpendicular to the TOF axis on an every-other-shot basis. The intensity difference of the parallel and the perpendicularly polarized photolysis beam was measured to be less than 0.2%. The TOF profiles taken with parallel photolysis polarization \( (I_p) \) and perpendicular polarization \( (I_l) \) were separately averaged. The isotropic \( (I_{ISO} = I_p + 2I_l) \) and anisotropic \( (I_{ANISO} = 2(I_l - I_p)) \) components of the TOF profiles were used to extract the speed dependent spatial anisotropy of the CH₃ products by fitting these components of the TOF profiles to basis functions generated by Monte Carlo simulations.

Spatial anisotropies of BrCl \( (\beta_{\text{phot}}) \) at the wavelengths 410, 420, and 430 nm that are needed for the analysis and the interpretation of data were obtained by a separate experiment. The Cl\( (^2P_{3/2}) \) and Cl\( (^2P_{1/2}) \) photofragments were ionized via 2 + 1 REMPI within 10 ns of the photolysis pulse, and the isotropic and anisotropic components of TOF profiles were analyzed to obtain the \( \beta_{\text{phot}} \) parameters.

### III. RESULTS AND DISCUSSION

#### A. Spatial anisotropy of BrCl

Table I lists the measured spatial anisotropies \( (\beta_{\text{phot}}) \) at the wavelengths 410, 420, and 430 nm. The Cl\( ^* \) channel shows strong parallel character \( (\beta \approx +2) \), whereas Cl channel shows perpendicular character \( (\beta \approx -1) \), which slightly changes across the wavelengths used. To check for possible orbital alignment\(^2\) of the Cl photofragments \( (\text{Cl}^* \text{ atoms cannot show alignment}) \) that may affect our spatial anisotropy measurement, we varied the direction of polarization of probe beam, and obtained the same values within the error bars. Therefore, the measured spatial anisotropies do not suffer from the interference caused by the orbital alignment of chlorine atoms. The contribution from “hot-band” dissociation from BrCl \( (\nu'' = 1 \text{ or } 2) \rightarrow \text{Br} + \text{Cl} + \text{Cl}\( ^* \))\(^2\) is found to be negligible at these wavelengths based on our analysis of TOF profiles. Our results show reasonable agreement with the results of Cooper et al.,\(^2\) although slight systematic differences exist, especially for the Br + Cl\( ^* \) channel.

### B. Speed distribution and spatial anisotropy of the CH₃ product

Figure 2 shows the isotropic and anisotropic components of representative CH\( _3 (\nu = 0) \) TOF profiles (via the Q branch of the band) at the collision energy \( E_{\text{coll}} = 0.14 \text{ eV} (\lambda_{\text{phot}} = 420 \text{ nm}) \), along with the fit to the basis functions. Overall shapes of the TOF profiles are very similar to the profiles obtained by Kandel and Zare\(^2\) with pure Cl\( (^2P_{3/2}) \) ground-state reaction at a similar collision energy. We were not able to detect CH\( _3 (\nu = 1) \) product via the 2 \( ^1 \) band within our signal-to-noise ratio, which gives an upper bound for the production of CH\( _3 (\nu = 1) \) of \( \leq 1\% \) relative to CH\( _3 (\nu = 0) \). The results of the fit give the laboratory speed distribution and the speed-dependent spatial anisotropy \( \beta_{\text{rxn}} (\nu) \) of CH\( _3 (\nu = 0) \) product, which are shown in Fig. 3.

The physically allowed speed of the CH\( _3 \) product is determined by the kinematics of the reaction, and the shape of the distribution within the allowed speed range is dependent on the differential cross section of the reaction product. As seen in Fig. 3(A), a significant portion of the distribution lies outside of the speed range that is allowed for Cl + CH\( _4 (\nu = 0) \) reaction at \( E_{\text{coll}} = 0.14 \text{ eV} \), which indicates that one or more other reaction channels contribute. Figure 3(B) shows the spatial anisotropy of the reaction product, \( \beta_{\text{rxn}} (\nu) \) for the CH\( _3 \) at \( E_{\text{coll}} = 0.14 \text{ eV} \), where the Cl\( ^* \) yield is 50\%. The measured spatial anisotropy is plotted along with four calculated curves. These show the predicted \( \beta_{\text{rxn}} (\nu) \) values for the Cl + CH\( _4 (\nu = 0) \) [curve (A), thin solid line], Cl + CH\( _4 [\nu = (\text{bending}) = 1] \) [curve (B), thick solid line], Cl + CH\( _4 [\nu = (\text{torsion}) = 1] \) [curve (C), dashed-dotted line], and Cl\( ^* + \text{CH}_4 (\nu = 0) \) [curve (D), dotted line]. The spatial anisotropy could be extracted reliably within the speed range that covers 99.5\% of the CH\( _3 (\nu = 0) \) speed distribution. Therefore, the trends in measured spatial anisotropy reflect the behavior of almost all of the CH\( _3 \) products. The measured spatial anisotropy does not have any similarity to the curve (D) nor can any combination of curves that include curve (D)

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**TABLE I.** Spatial anisotropies of BrCl photolysis and the relative yields of Cl\( ^* \).

<table>
<thead>
<tr>
<th>( \lambda_{\text{phot}} ) (nm)</th>
<th>( E_{\text{coll}} ) (eV)</th>
<th>%Cl( ^* )</th>
<th>( \beta_{\text{phot}} ) (Cl)</th>
<th>( \beta_{\text{phot}} ) (Cl( ^* ))</th>
</tr>
</thead>
<tbody>
<tr>
<td>410</td>
<td>0.13</td>
<td>37</td>
<td>(-0.76\pm0.02)^c</td>
<td>1.88\pm0.02)^c</td>
</tr>
<tr>
<td>420</td>
<td>0.14</td>
<td>50(^b)</td>
<td>(-0.69\pm0.01)</td>
<td>1.79\pm0.04</td>
</tr>
<tr>
<td>430</td>
<td>0.16</td>
<td>57</td>
<td>(-0.47\pm0.02)</td>
<td>1.82\pm0.04</td>
</tr>
</tbody>
</table>

\(^a\) See Cao et al.\(^19\) See also Cooper et al.\(^20\)

\(^b\) Value at 421.7 nm.

\(^c\) Uncertainties shown are 1 \( \sigma \).
provide a reasonable fit to the measured anisotropy. Instead, curves (B) and (C), corresponding to the reaction with \( \nu_4 = 1 \) and \( \nu_2 = 1 \) vibrationally excited methane, respectively, provide an exceedingly good match to the data, with curve (B) giving a marginally better agreement than curve (C) to the measured anisotropy. This result indicates that the previously proposed Cl* reactivity with CH\(_4\) (\( \nu = 0 \)) is unimportant compared with ground-state Cl reactivity with CH\(_4\) (\( \nu_2 \) or \( \nu_4 = 1 \)). The latter accounts for only \( \sim 1\% \) of the CH\(_4\) population in our beam expansion. Similar behavior, the enhanced reactivity of Cl(\( ^2P_{3/2} \)) + CH\(_4\) (\( \nu_2 \) or \( \nu_4 = 1 \)) over Cl(\( ^2P_{3/2} \)) + CH\(_4\) (\( \nu = 0 \)) was first observed by Kandel and Zare.\(^21\)

Although the Cl* reaction channel is exothermic, it is possible that the reaction barrier might be higher than that of the Cl–atom reaction owing to nonadiabatic interaction near the barrier. Clearly, the barrier height for the Cl* reaction is not smaller than that of the Cl–atom reaction as measured from the energy level of Cl(\( ^2P_{3/2} \)) + CH\(_4\). Therefore, it is possible that the collision energy of 0.14 eV might not be sufficient to surmount the Cl* reaction barrier. On the other hand, Cl* reactivity would be most noticeable at lower collision energies, where the total energies are closer or below reaction threshold for Cl + CH\(_4\) (\( \nu = 0 \)). To check for these possibilities, we tuned the 0.14 eV collision energy to higher and lower values. Experimentally, tuning of the collision energy is limited by two factors. The Cl* branching ratio of BrCl photolysis drops sharply at photolysis wavelengths below 410 nm.\(^19\) At photolysis wavelengths longer than 430 nm, the CH\(_3\) signal level drops significantly such that it does not allow us to obtain a reliable spatial anisotropy of the product. In Fig. 4, we compare the results of the analysis at \( E_{\text{coll}} = 0.13, 0.14, \) and 0.16 eV, with the calculated anisotropies as described in the previous paragraph. At \( E_{\text{coll}} = 0.13 \) eV (\( \lambda_{\text{phot}} = 430 \) nm), an appreciable portion of the Cl + CH\(_4\) (\( \nu = 0 \)) collision energy distribution has lower energy than the activation energy of 960 cm\(^{-1}\) (see Fig. 1). Again, all of the measured spatial anisotropies lie close to the curves expected for Cl + CH\(_4\) (\( \nu_2 \) or \( \nu_4 = 1 \)) \( \rightarrow \) HCl + CH\(_3\). Therefore, we conclude that the Cl* + CH\(_4\) reaction is not important in the collision energy range between 0.13 and 0.16 eV; instead, we suggest the dominant role of reaction of methane molecules excited with one quantum of vibration in \( \nu_2 \) or \( \nu_4 \) with ground-state Cl atoms. Although we cannot claim that the Cl* + CH\(_4\) (\( \nu = 0 \)) channel is less reactive than Cl + CH\(_4\) (\( \nu = 0 \)), we can safely state that the Cl* + CH\(_4\) (\( \nu = 0 \)) reaction is less reactive than the Cl + CH\(_4\) (\( \nu_2 \) or \( \nu_4 = 1 \)) reaction. Moreover, we can set a conservative upper bound for the Cl* reactivity to be \( \leq 15\% \) of the reactivity of Cl + CH\(_4\) (\( \nu_2 \) or \( \nu_4 = 1 \)) reaction near reaction threshold.

In 1972, Truhlar\(^26\) and Muckerman and Newton\(^27\) first discussed whether the mechanism for the F* + H\(_2\) \( \rightarrow \) HF + H reaction channel should be described as statistical vs adiabatic. In the statistical limit, each spin-orbit state contributes according to its multiplicity to the overall cumulative probability, thus predicting appreciable contributions from the spin-orbit excited halogen atom. In the adiabatic limit, no F* reactivity is predicted to generate HF(\( ^1\Sigma^+ \)) + H products. Schatz\(^28\) in his theoretical work on Cl(\( ^2P_{3/2} \)) + HCl \( \rightarrow \) HCl + Cl(\( ^2P_{1/2} \)) (where \( J, J' = 1/2 \) or 3/2) reaction, generalized statistical and adiabatic approaches. Although

![FIG. 3. (A) Speed distribution of CH\(_3\) product (open circles). The scale bar on the top shows the laboratory frame speed range allowed for Cl+CH\(_4\) (\( \nu = 0 \)) \( \rightarrow \) HCl + CH\(_3\). (B) Speed-dependent spatial anisotropy \( \beta_{\text{rad}}(\nu) \) for CH\(_3\) product 0.14 eV collision energy (open circles). The experimental data are presented along with four curves showing the predicted spatial anisotropy for the Cl + CH\(_4\) (\( \nu = 0 \)) (thin solid line), Cl + CH\(_4\) (\( \nu_2 = 1 \)) (thick solid line), Cl + CH\(_4\) (\( \nu_4 = 1 \)) (dashed-dotted line), and Cl* + CH\(_4\) (\( \nu = 0 \)) (dotted line). The data indicate that reaction with \( \nu_2 \) or \( \nu_4 = 1 \) excited methane is dominant. Error bars represent 2 \( \sigma \) of multiple (at least five) sets of measurements. The range of the \( y \) axis is chosen such that the minimum value corresponds to \( \beta_{\text{rad}}(\text{Cl}) \), and the maximum value to \( \beta_{\text{rad}}(\text{Cl}^*) \).](https://image.pollinations.ai/prompt/FIG. 3. (A) Speed distribution of CH\(_3\) product (open circles). The scale bar on the top shows the laboratory frame speed range allowed for Cl+CH\(_4\) (\( \nu = 0 \)) \( \rightarrow \) HCl + CH\(_3\). (B) Speed-dependent spatial anisotropy \( \beta_{\text{rad}}(\nu) \) for CH\(_3\) product 0.14 eV collision energy (open circles). The experimental data are presented along with four curves showing the predicted spatial anisotropy for the Cl + CH\(_4\) (\( \nu = 0 \)) (thin solid line), Cl + CH\(_4\) (\( \nu_2 = 1 \)) (thick solid line), Cl + CH\(_4\) (\( \nu_4 = 1 \)) (dashed-dotted line), and Cl* + CH\(_4\) (\( \nu = 0 \)) (dotted line). The data indicate that reaction with \( \nu_2 \) or \( \nu_4 = 1 \) excited methane is dominant. Error bars represent 2 \( \sigma \) of multiple (at least five) sets of measurements. The range of the \( y \) axis is chosen such that the minimum value corresponds to \( \beta_{\text{rad}}(\text{Cl}) \), and the maximum value to \( \beta_{\text{rad}}(\text{Cl}^*) \).)

![FIG. 4. Spatial anisotropies (open circles) for CH\(_3\) products at collision energies \( E_{\text{coll}} = 0.16 \) (top), 0.14 (middle), and 0.13 (bottom) eV. Also shown are Cl + CH\(_4\) (\( \nu = 0 \)) (thin solid line), Cl + CH\(_4\) (\( \nu_2 = 1 \)) (thick solid line), and Cl* + CH\(_4\) (\( \nu = 0 \)) (dotted line).](https://image.pollinations.ai/prompt/FIG. 4. Spatial anisotropies (open circles) for CH\(_3\) products at collision energies \( E_{\text{coll}} = 0.16 \) (top), 0.14 (middle), and 0.13 (bottom) eV. Also shown are Cl + CH\(_4\) (\( \nu = 0 \)) (thin solid line), Cl + CH\(_4\) (\( \nu_2 = 1 \)) (thick solid line), and Cl* + CH\(_4\) (\( \nu = 0 \)) (dotted line).)
the calculated overall rate constant is shown to be rather insensitive to the degree of nonadiabatic interaction for this particular reaction, Schatz claimed that this system follows the adiabatic limit for a physically appropriate value of spin-orbit coupling for the Cl atom. In sharp contrast, several rate constant calculations on the Cl + CH₄ reaction¹²,¹³ have been based on a statistical treatment of the Cl* reaction channel, including significant contributions from the Cl⁺ reaction channel.

Kandel and Zare²¹ estimated the low-frequency torsion (ν₂) or bending mode (ν₄) enhancement to be 200 times compared to the Cl+CH₄(ν=0) reaction, which is larger than the enhancement factor of 30 for the asymmetric stretching mode excited (ν₃=1)CH₄ reaction.¹⁷ At this point, it is not clear why the bending mode or the torsion mode excitation shows more enhancement than the asymmetric stretching mode, recalling that all theoretical calculations¹²-¹⁵ support a collinear Cl–H–C transition-state geometry. Kandel and Zare also suggested that the bending or torsion mode enhancement might contribute to the non-Arrhenius behavior at low temperatures. Theoretical studies by Duncan and Truong¹⁵ and Corchado et al.¹² predict modest to large enhancement of reactivity by bending mode excitation caused by the lowering or removal of the barrier to reaction. Their calculations, at the same time, suggest that the major contribution to the non-Arrhenius behavior is quantum mechanical tunneling rather than bending mode enhancement. Recent work by Michelsen and Simpson² suggest a modest contribution of bending enhancement to the non-Arrhenius behavior, with the major contribution from tunneling.

Our results set a new upper bound for Cl⁺ reactivity with CH₄, and they rule out the possibility of the role of Cl⁺ reactivity in the explanation of the non-Arrhenius behavior at low temperatures. Also, we observed the dominant role of reaction of ground state chlorine atom with bending or torsion mode excited methane at near-threshold collision energies. Further detailed studies may reveal the nature of the bending and torsion mode enhancement in the Cl+CH₄ reaction.

ACKNOWLEDGMENTS

The authors thank Andrew J. Orr-Ewing, S. Alex Kandel, and Hope A. Michelsen for valuable discussions and Andrew J. Orr-Ewing for providing them with unpublished data on BrCl photofragment anisotropies. One of the authors (H.A.B.) also thanks the National Science Foundation for a graduate fellowship. This work was supported by the National Science Foundation under Grant No. NSF-CHE-99-00305.

²⁴A. J. Orr-Ewing (private communications, 2001): The β parameters at longer wavelengths 460–560 nm depend very strongly on the vibrational quantum number of the BrCl parent. This “hot-band” (ν=1 or 2) photodissociation is observed to give different β parameters from those of BrCl (ν=0) photolysis. This is also observed for BrF and it is a consequence of Franck–Condon factors for excitation to different repulsive surfaces. See also E. Wrede, S. Laubach, S. Schulenburg, A. Brown, E. R. Wouters, A. J. Orr-Ewing, and M. N. R. Ashfold, J. Chem. Phys. 114, 2629 (2001).
²⁵If we assume that CH₃ product distributions from Cl and Cl⁺ reactions give the same shape of speed distribution, the effective spatial anisotropy βₛᵥ(x) is given by the reactivity and flux weighted average of two reactions: βₛᵥ(x)=(1−f)βᵣ(x)+fβ⁺(x), where βᵣ and β⁺ are the calculated spatial anisotropies for Cl+CH₄ (νᵣ=0) and Cl⁺+CH₄ (ν₊=0) reactions, and f=ω₃⁺[Cl⁺][CH₄(νᵣ=0)]/ω₃[Cl⁺][CH₄(ν₊=0)] (0.99 for Eᵣ=0.14 eV). The w is the relative reaction cross section of Cl⁺ atoms, and ω and a are center of mass speeds associated with Cl and Cl⁺, respectively. The w is increased from zero until the βₛᵥ(x) curve lies completely outside of the 1 σ of the measured spatial anisotropy for all speeds. We chose w as the upper bound for the Cl⁺ reactivity.